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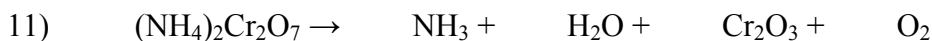
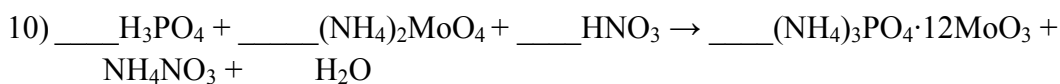
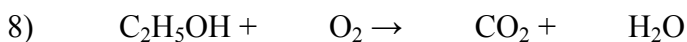
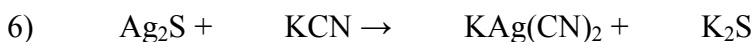
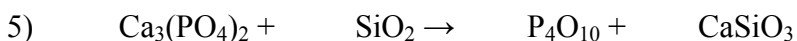
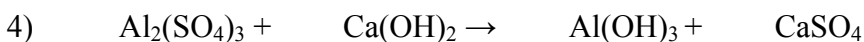
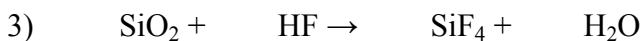
AP Chemistry

Balancing Equations Review

Example:

Chromium compounds exhibit a variety of bright colors. When solid ammonium dichromate, $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$, a vivid orange compound is ignited, a spectacular reaction occurs. The products are solid chromium (III) oxide, nitrogen gas, and water vapor. Balance the equation for this reaction.

Questions:



- 12) The combustion of ethanol ($\text{C}_2\text{H}_5\text{OH}$) forms carbon dioxide and water vapor. A combustion reaction refers to a reaction of a substance with oxygen gas.
- 13) Aqueous solution of lead (II) nitrate and sodium phosphate are mixed, resulting in the precipitate formation of lead (II) phosphate with aqueous sodium nitrate as the other product
- 14) Solid zinc reacts with aqueous HCl to form aqueous zinc chloride and hydrogen gas.
- 15) Aqueous strontium hydroxide reacts with aqueous hydrobromic acid to produce water and aqueous strontium bromide.
- 16) Solid calcium carbide (CaC_2) reacts with water to form an aqueous solution of calcium hydroxide and acetylene gas, C_2H_2 .
- 17) When solid potassium chlorate is heated, it decomposes to form solid potassium chloride and oxygen gas.
- 18) Solid zinc metal reacts with sulfuric acid to form hydrogen gas and an aqueous solution zinc sulfate.
- 19) When liquid phosphorus trichloride is added to water, it reacts to form aqueous phosphorus acid, H_3PO_4 , and aqueous hydrochloric acid.
- 20) When hydrogen sulfide gas is passed over solid hot iron (III) hydroxide, the resultant reaction produces solid iron (III) sulfide and gaseous water.